Arsenic(III)

Equilibrium reaction	lg$K$ at infinite dilution and T = 298 K		
	Baes and Mesmer, 1976	Nordstrom and Archer, 2003	Nordstrom et al., 2014
As(OH)$_4^-$ + H$^+$ $\rightleftharpoons$ As(OH)$_3$ + H$_2$O		9.17	9.24 ± 0.02

C.F. Baes and R.E. Mesmer, The Hydrolysis of Cations. Wiley, New York, 1976, p. 375.

D.K. Nordstrom and D. Archer, Arsenic thermodynamic data and environmental geochemistry. In: Arsenic in Ground Water. Welch AH, Stollenwerk KG (eds) Kluwer Academic Publishers, Amsterdam, 2003, pp. 1–25.

D.K. Nordstrom, J. Majzlan and E. Königsberger, Thermodynamic properties for As minerals & aqueous species. Reviews in Mineralogy & Geochemistry, 79, 217–255 (2014).

Distribution diagrams

These diagrams have been computed at two As(III) concentrations (1 mM = 1×10^{-3} mol L^{-1} and 1 µM = 1×10^{-6} mol L^{-1}) with the 'best' equilibrium constant above (in green). Calculations assume T = 298 K for the limiting case of zero ionic strength (*i.e.*, even neglecting plotted ions).