Arsenic(V)

Equilibrium reaction	lg*K* at infinite dilution and *T* = 298 K		
	Khodakovsky et al. (1968)	Nordstrom and Archer, 2003	Nordstrom et al., 2014
$H_2AsO_4^- + H^+ \rightleftharpoons H_3AsO_4$	2.21	2.26 ± 0.078	2.25 ± 0.04
$HAsO_4^{2-} + H^+ \rightleftharpoons H_2AsO_4^-$	6.93	6.99 ± 0.1	6.98 ± 0.11
$AsO_4^{3-} + H^+ \rightleftharpoons HAsO_4^{2-}$	11.51	11.80 ± 0.1	11.58 ± 0.05

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D.K. Nordstrom, J. Majzlan and E. Königsberger, Thermodynamic properties for As minerals & aqueous species. Reviews in Mineralogy & Geochemistry, 79, 217–255 (2014).

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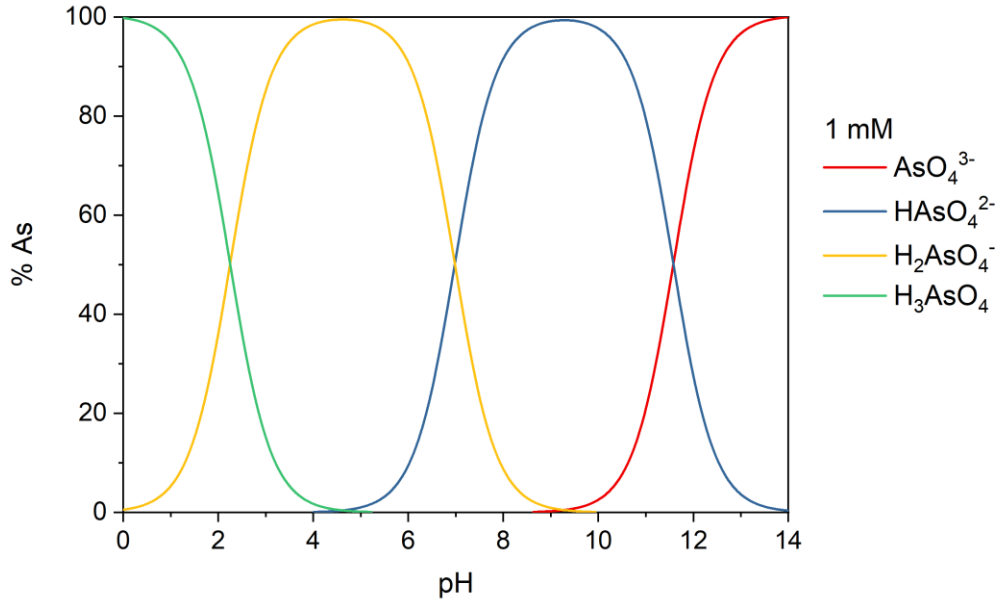
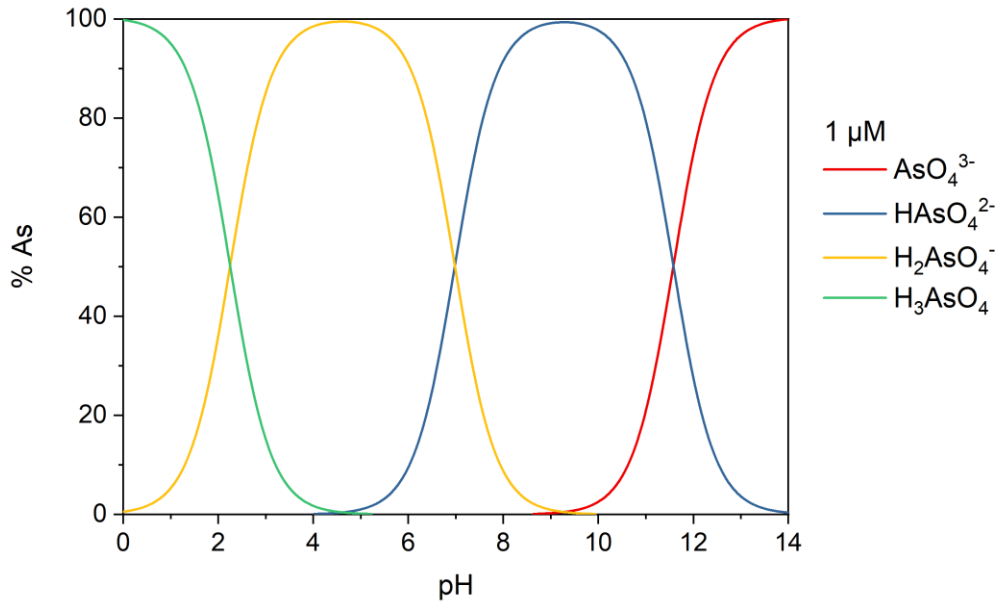
Distribution diagrams

These diagrams have been computed at two As(V) concentrations (1 mM = $1x10^{-3}$ mol L^{-1} and 1 µM = $1x10^{-6}$ mol L^{-1}) with the 'best' equilibrium constants above (in green). Calculations assume T = 298 K for the limiting case of zero ionic strength (*i.e.*, even neglecting plotted ions).





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