

Copper(I)

Equilibrium reactions	lgK at infinite dilution and $T = 298 \text{ K}$
	Brown and Ekberg, 2016
$\text{Cu}^+ + \text{H}_2\text{O} \rightleftharpoons \text{CuOH} + \text{H}^+$	-7.8 ± 0.4
$\text{Cu}^+ + 2 \text{H}_2\text{O} \rightleftharpoons \text{Cu}(\text{OH})_2^- + 2 \text{H}^+$	-18.6 ± 0.6

P.L. Brown and C. Ekberg, Hydrolysis of Metal Ions. Wiley, 2016, pp. 650–702.

Distribution diagrams

These diagrams have been computed at two Cu(I) concentrations ($1 \text{ mM} = 1 \times 10^{-3} \text{ mol L}^{-1}$ and $1 \text{ }\mu\text{M} = 1 \times 10^{-6} \text{ mol L}^{-1}$) with the 'best' equilibrium constants above. Calculations assume $T = 298 \text{ K}$ for the limiting case of zero ionic strength (*i.e.*, even neglecting plotted ions).

