Lead(IV)

Equilibrium reactions	lg*K* at infinite dilution and *T* = 298 K
	Feitknecht and Schindler, 1963
$\beta\text{-}PbO_2 + 2\ H_2O \rightleftharpoons Pb^{4+} + 4\ OH^-$	–64
$\beta\text{-}PbO_2 + 2\ H_2O + 2\ OH^- \rightleftharpoons Pb(OH)_6^{2-}$	–4.5

W. Feitknecht and P. Schindler, Solubility constants of metal oxides, metal hydroxides and metal hydroxide salts in aqueous solution. Pure Appl. Chem., 6, 125–206 (1963).

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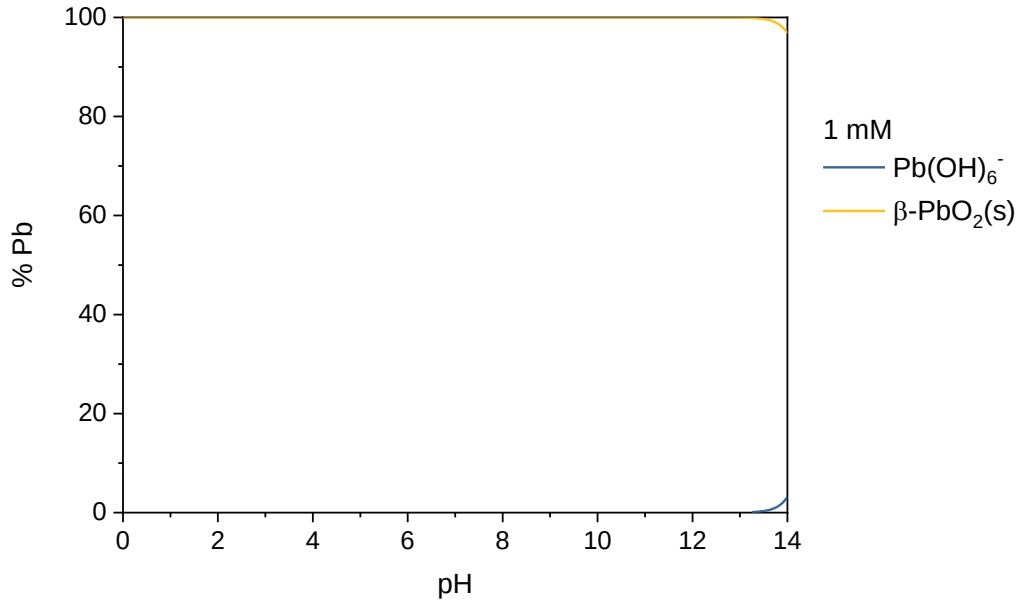
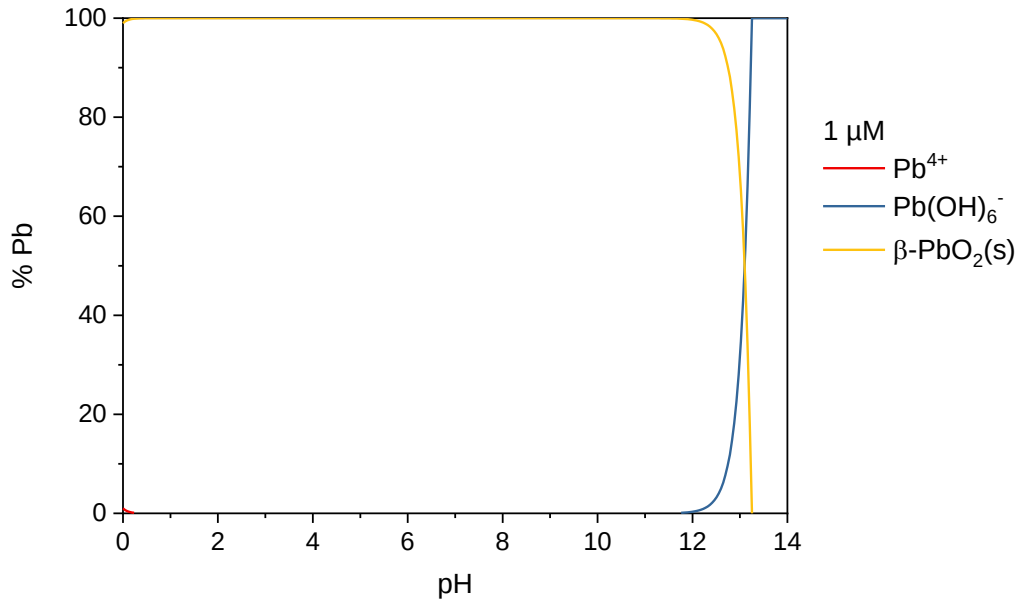
Distribution diagrams

These diagrams have been computed at two Pb(IV) concentrations (1 mM = $1x10^{-3}$ mol L^{-1} and 1 μM = $1x10^{-6}$ mol L^{-1}) with the 'best' equilibrium constants above. Calculations assume T = 298 K for the limiting case of zero ionic strength (*i.e.*, even neglecting plotted ions).





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