Tellurium(-II)

Equilibrium reactions	lg*K* at infinite dilution and *T* = 298 K
	Filella and May, 2019[a]
$Te^{2-} + H^+ \rightleftharpoons HTe^-$	11.81
$HTe^- + H^+ \rightleftharpoons H_2Te$	2.476

[a]The number of significant figures are retained to minimise propagation of round-off errors; they should not be taken to indicate the relative uncertainty of the values, which is always at least one order of magnitude less than indicated.

M. Filella and P.M. May, The aqueous chemistry of tellurium: critically-selected equilibrium constants for the low-molecular-weight inorganic species. Environ. Chem. 16, 289–295 (2019). doi:10.1071/EN19017

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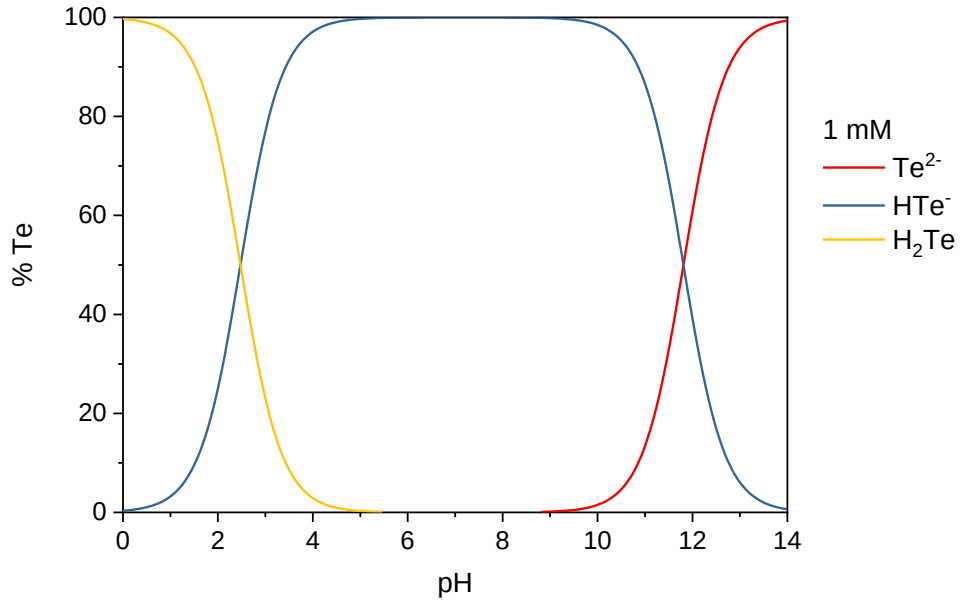
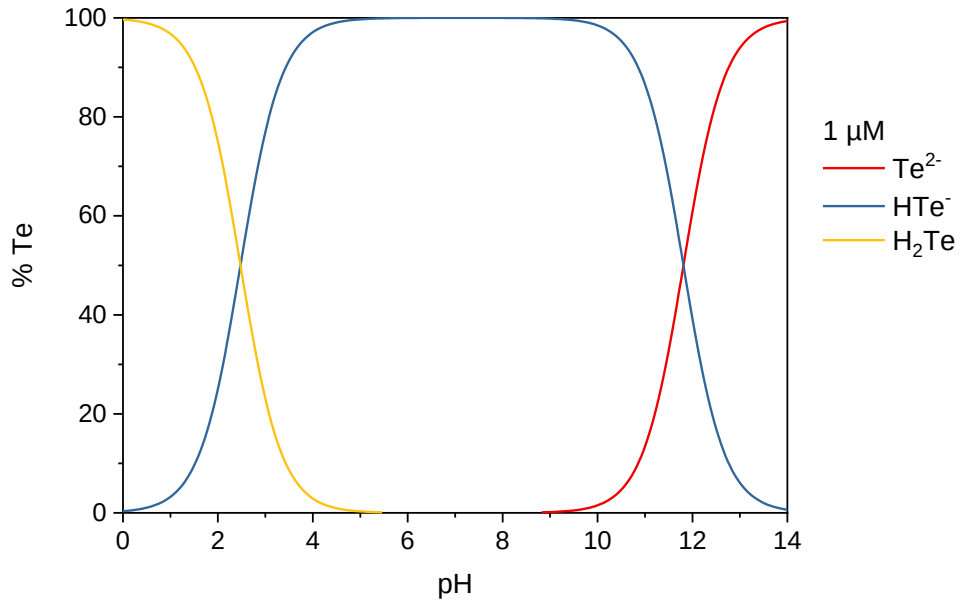
Distribution diagrams

These diagrams have been computed at two Te(-II) concentrations (1 mM = 1x10^{-3} mol L^{-1} and 1 µM = 1x10^{-6} mol L^{-1}) with the 'best' equilibrium constants above. Calculations assume T = 298 K for the limiting case of zero ionic strength (*i.e.*, even neglecting plotted ions).





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